

## Chapter 6 Chemical Bonding Section 2 Covalent Answer Key

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[Chapter 6 Chemical Bonding Section](#)

CHAPTER 6 REVIEW Chemical Bonding SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2. b A covalent bond consists of (a) a shared electron.

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Chapter 6: Chemical Bonding Section 6-3: Ionic Bonding and Ionic Compounds \_\_\_\_\_ Pacing Regular Schedule: with lab(s): 3 days without lab(s): 1 day Block Schedule: with lab(s): 1 1/2 days without lab(s): 1/2 day Objectives 1. Compare and contrast a chemical formula for a molecular compound with one for an ionic compound. 2.

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CHAPTER 6 REVIEW Chemical Bonding SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. a The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom. 2. d In a crystal of an ionic compound, each cation is surrounded by a number of (a) molecules. (c) dipoles. (b) positive ions. (d) negative ions. 3. b Compared with the ...

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Chapter 6: Chemical Bonding Section 6-4: Metallic Bonding \_\_\_\_\_ Pacing Regular Schedule: with lab(s): 3 days without lab(s): 1 day Block Schedule: with lab(s): 1 1/2 days without lab(s): 1/2 day Objectives 1. Describe the electron-sea model of metallic bonding, and explain why metals are good electrical conductors. 2. Explain why metal surfaces ...

[Chapter 6: Chemical Bonding](#)

Chapter 6: Chemical Bonding Section 2: Covalent Bonding and Molecular Compounds Overview We will define molecule and molecular formula. We will explain the relationships between potential energy, distance between approaching atoms, bond length, and bond energy. We will state the octet rule. We will see how to write Lewis structures.

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Section 6.1 – Ionic Bonding (pages 158-164) Key Concepts: When is an atom unlikely to react? What is one way in which elements can achieve stable electron configurations? How does the structure of an ionic compound affect its properties? Vocabulary: electron dot diagram. ion. anion. cation. chemical bond. ionic bond. chemical formula. crystals

[Chapter 6 – Chemical Bonds \(Pages 156-175\)](#)

This course is an OER section developed by Dr. Ara Kahyaoglu for Bergen Community College. The primary text was developed for the Saylor Academy and is modified to better serve the course objectives for BCC students.Chapter 6 - Chemical Bonding6.1 Lewis Electron Dot DiagramsLearning Objectives 1. Draw a Lewis electron dot diagram for an atom or a monatomic ion.6.2 Electron Transfer: Ionic ...

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Section 2 Covalent Bonding and Molecular Compounds. Molecule – compound held together by covalent bonds. Chemical formula – give the atoms and the number of atoms that make up a compound. Bond energy is the energy required to break a chemical bond. The Octet Rule – Compounds gain, lose, or share electrons to have 8 valence electrons ...

[Chapter 6 Chemical Bonding Section 1 Introduction to ...](#)

Chapter 6: Chemical Bonding Section 1- Introduction to Chemical Bonding Objectives: define chemical bond; differentiate between covalent and ionic bonding; explain why bonding occurs; use the difference in electronegativity to determine whether a bond is polar covalent, nonpolar covalent or ionic

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CHAPTER 6 REVIEW . Chemical Bonding . SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_\_;b\_\_ In metals, the valence electrons are considered to be (a) attached to particular positive ions. (c) immobile. (b) shared by all surrounding atoms. (d) involved in covalent bonds. 2.

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Chapter 6: The Structure of Matter. Section 2: Ionic and Covalent Bonding. Objectives. Why do atoms form bonds? ... There are two basic kinds of chemical bonding: ionic bonding . covalent bonding. Ionic Bonds. ... Chapter 6: The Structure of Matter

[Chapter 6: The Structure of Matter](#)

Chapter 6 Chemical Bonds Section 6.2 Covalent Bonding (pages 165–169) This section discusses the formation of covalent bonds and the factors that determine whether a molecule is polar or nonpolar. It also discusses attractions between molecules. Reading Strategy(page 165) Relating Text and Visuals As you read the section, look closely at ...

[Chapter 6 Chemical Bonds Section 6.2 Covalent Bonding](#)

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